

Chemistry

Standard #10- Convert number of particles, grams, and liters of gas to moles and calculate the percent composition of an element in a compound.

Chapter 10: Chemical Quantities

- Measuring matter- grams and moles
- Mole-mass-volume relationships
- STP
- Percent composition
- Empirical formula vs molecular formula

Journal Work

- Define the **vocabulary terms** from the chapter in your journal.
- Read each section and answer the **Section Assessment** and **Chapter Assessment** questions in the chapter (#1-105) odds in your journal. Self-check answers with key in the back of book.
- Answer the following questions in hand-written **paragraph** format in journal
 - What is the historical significance of the mole concept and how was it determined?
- Read and Summarize a current article related to one of the major topics found in the chapter. Attach a copy of the article to your written summary in your journal.

Assessment	Teacher Initials	Score	Date
1. Show completed journal . <ul style="list-style-type: none"> ○ <i>Vocabulary</i> ○ <i>Questions</i> ○ <i>Paragraphs</i> ○ <i>Article Summary</i> 			
2. Take and pass the exam .			
3. After passing the exam, do the lab work Lab 9: Molecular mass of a gas (handout)			

You have completed Standard #10

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